
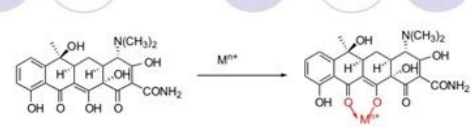


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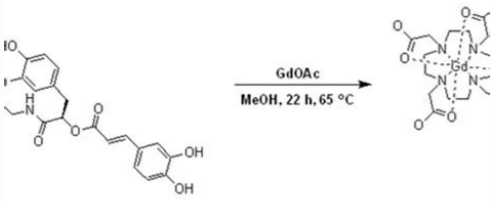
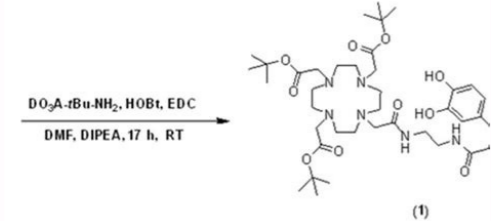
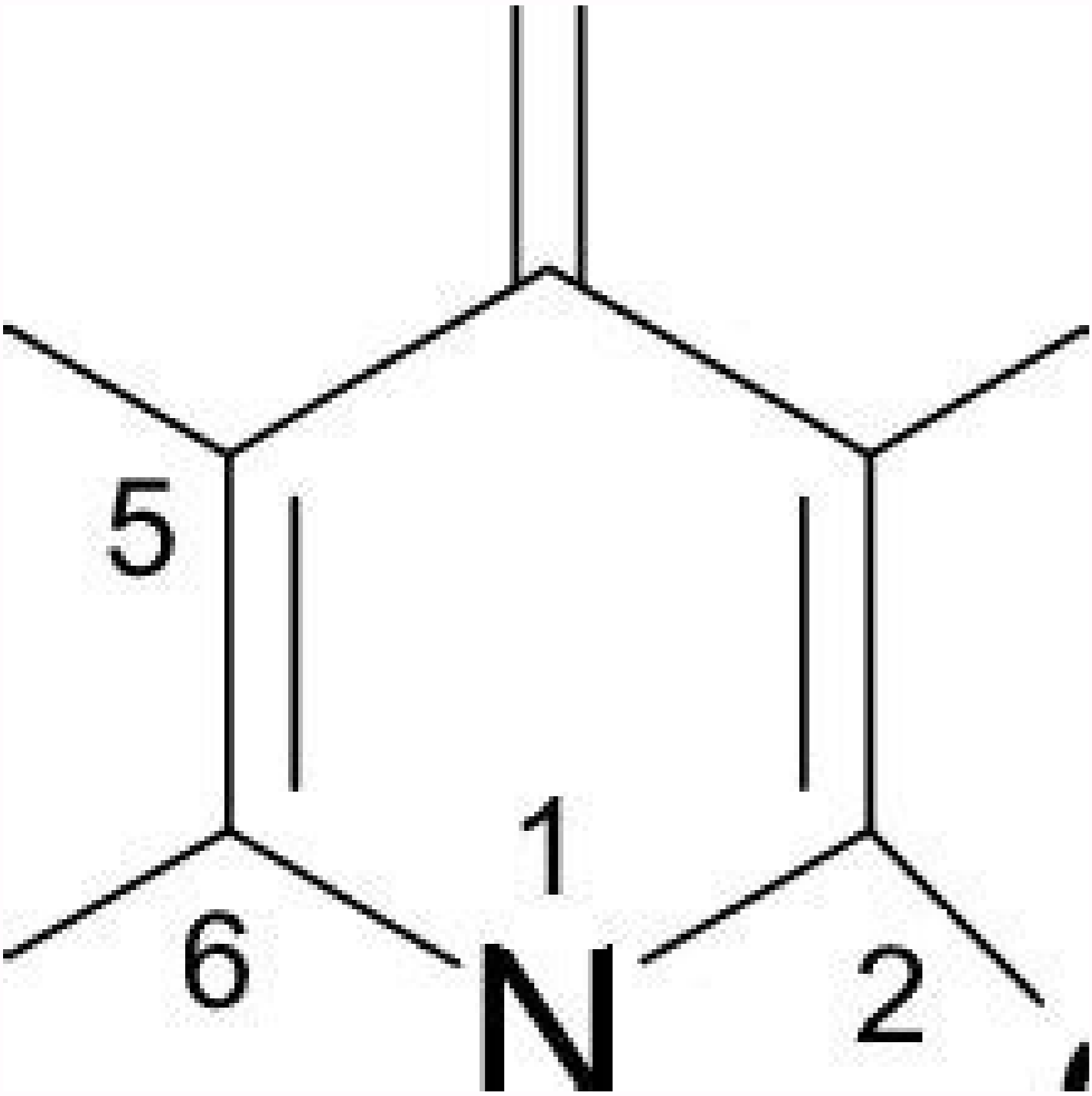
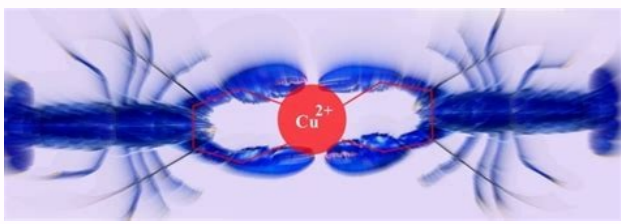
# Chelate complex formation

## Formation of metal chelates



- Stable chelate complexes are formed by the tetracycline with many metals, including calcium, magnesium and iron. Such chelates are usually insoluble in water.
- The tetracyclines are distributed into the milk of lactating mothers and will cross the placental barrier into the fetus.

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In addition, these adventitious chelators, several biomolecules occur specifically to join certain metals. Let's take a look at an example: copper (II) affinities for ethylenediamine (en) and methylamine are compared.  $[Cu^{2+} + 2MenH_2] \rightleftharpoons [Cu(en)]^{2+} + 2MenH_2$  \ The copper ion forms a complex of ethylenediamine chelate in the first equation. The chelating agents, unlike the other ligands in the coordination compounds, are joined through the multiple and themes in the mole of the ligand, not only one. For example, it is known that *Pseudomonas* species secrete pyocyanin and pyoverdinin that bind to iron. Quelacion is a very common tune used in different branches of science such as chemical, biology and more sciences. Between a pointed ligand and a single central thing, two or more different links of different coordinates are formed or present. These agents derive from microorganisms. The work results in the creation of a  $[Cu(en)]^{2+}$  chelature ring with five members. Prismaticly all biochemics exhibit the ability to dissolve certain metallic cations. It is difficult to explain the thermodynamic values in terms of changes in the molecular solution exactly, but it is evident that the effect is mainly an entropy effect. Quelato complex The ligands (electron donors) used in the process of quelacion are known as chelants, chelars and kidnappers. Quelatecécical Metal-Edta structure of the EDTA. Quelat complexes are contrasted with coordination complexes composed of monodentated ligands, which form only a link with the central time. Enterobactin, produced by *E. Coli* (en) concentration is much greater than concentration  $[Cu(en)]^{2+}$  because the analytical concentration of methylation and the concentration of the concentration of copper is the same in both reactions. As we know,  $\Delta G = \Delta H - T\Delta S$  " Among the two stability constants is due to the effects of entropy, since the enthalpy must be almost the same for both reactions. Although they can be beneficial in cases of heavy metal poisoning, agents that can also be dangerous. The work results in the formation of a five -member ring. Quelacion is the formation or presence of two or more separate coordinate links between a pooling ligand (mostleuple union) and a single central thing. The effect increases with the number of chelato rings, so the concentration of the EDTA complex, which has six chelature rings, is much greater than a corresponding complex with two monodentate nitrogen donors and four monodentated carboxylate ligands. Many microbial species produce water soluble pigments that serve as chelating agents, called sideroforos. Therefore, the phenomenon of the chelature effect is a firmly established empirical fact. This therapy was approved by the Food and Medicines Administration of the USA. In 1991. Examples. Explain the process of cheating with an example. Give some uses of chelating agents. How does the Quelat effect take place? Draw an ethylenediamine structure and explain its uses. Do you know? Quelacion therapy sometimes causes fever and vintages in patients. Some of the chelating agents can cause respiratory failure. As there are cases of zinc and other inorganic molecules that are used as cheels. used in maintenance therapy to prevent copper absorption in people with Wilson's disease. Thequelization is ostile to provide nutritional supplements. It is used in the therapy of to eliminate the technical metals of the body. Used in homogosse manufacturing It is used in the treatment of chemical water to help the removal of metals and fertilizers. The cheating process is used by plants for the elimination of heavy metals. Agents of natural chelants and cosméc industry The skeleton agents are compounds that form multiple links with metallic ions. There are two particles to the left and one to the right in equation one, while there are three particles to the left and one to the right in equation two. This distinction suggests that when a chelato complex is formed with a bident ligand, less entropy the disorder is lost than when a complex with monodentated ligands is formed. In (2), the two monodentated methylamine ligands of approximately the same donor power (the enthalpy of link formation Cu. The applications of the chelants of the therapy of cheating are the use of chelating agents to detoxify the agents metallic poisonous lyric, such as mercury, arsã © nico and lead, by turning them into a chemically inert form that can be excreted without more interaction with the body. The effect of chelature describes the greatest affinity of the chelants for an ion for an ion methodical, in comparison with the affinity of a collection of similar (monodentated) similar ligands for the same metal. We go to our main question, what is the quelacion? which has two or more coordinated or dative bonds between a ligand (generally orgal) and a central metal enthusias Ormar links with metallic ions. Let's take a look at how the quelacion mechanism works. One of the variables that contribute to the Entropia gap is this. The bident ligand is replaced by two monodentated methylamine ligands with approximately the same resistance Donor in the second reaction, which suggests that the links of  $Cu^{2+} + n$  are similar in both reactions. The constant balance for for The reaction is taken into account in the thermodynamic approach to explain the work.  $[Cu^{2+}] + n [en] \rightleftharpoons [Cu(en)]^{2+}$  (1)  $[Cu^{2+}] + 2menH_2 \rightleftharpoons [Cu(Cu(Cu(Cu(Menh_2)_2)^{2+})]$  (2) In (1), the bident light ethylenediamine forms a chelature complex with the copper ion. The use of EDTA Disparte instead of Edta de Calcio has resulted in deaths due to hypocalcemia. Due to constant demand among consumers for using ecological products, natural chelating agents are now consumed on a large scale. The use of chelating agents helps Increase the ostile life of the products and make it less damage to the environment. Quelacion therapy is the use of chelating agents to detoxify poisonous metallic agents, such as mercury, arsyssal and lead, converts them into a very inert form that can be excreted without more interaction with the body. The process of cheating is widely used in the detoxify complex. Coli, is the most known quelantent agent. Organic compounds such as gly outteen and histidine aminoine, orgins rings as evil and polypys such as phytoquequelatina are also typical. In general, these ligands are organic compounds and are called chelants, cheers, chelants or kidnappers; The resulting complexes are called chelato compounds. Due to their organic nature, they are mainly used to produce cosmã © techniques. Such chelating agents include porphyrin rings in hemoglobin and chlorophyll. Consider the two balances, in aqueous solution, between the copper ion (II) (Cu<sup>2+</sup>) and ethylenediamine (en) on the one hand and methylamine (CH<sub>3</sub>NH<sub>2</sub> (menh<sub>2</sub>)) on the other. The changes of solveration and the formation of the chelato ring are two other factors to The changes of enthalp. Reason for the greatest stability of the Quelato complex. The greater the equilibrium constant, the greater the complex concentration. The clarity sake has eliminated eligric loads. Let us discuss the terms used in the definition of quelacion for a better understanding of definition. Lighting: a ligand is an ion or molecula that forms a coordination complex donating a couple of electrons to the central metal ion or ion. of the utomos used to join a central metal ion or ion vary between the controversial ligands. The work is backed by certain thermodynamic concepts. Hexadentated ligands, such as EDTA, have six donors with electron pairs that can join a central metal ion or ion. What is the chemistry in the chemical? Chelant ligands (monodentated). The compound, therefore, formed, is used in a variety of applications. Non -saturated oils can be obtained from these natural chelars when used with antioxidants such as tocopherol. Squares of the use of natural chelating agents: It helps to increase the ostile life of work products as a natural alternative to Edtaused for skin lighting functionality to make metallic ions inactive in the color. cil of using biodegradable forms and friendly friending of tóxic elements that are dedicated to natural animals such as microorganisms. Nucleic zeas are used to form links. The majority of the cullets dissolve the metallic cations to form complexes of cheatoethylandiamine, a bident ligand, forms a complex of chelature with copper. In conditions of equal copper concentrations and when the concentration of methylamine is double of ethylenediamine, the concentration of the complex in (1) will be greater than the concentration of the complex in (2). Natural chelating agents are biodegradable and free of tóxic elements. Subuses to stability constants show the stoichiometry of the complex, and square supports indicate the concentration. Quelacion is the formation or presence of two or more separate coordinate links between a pooling ligand (mostleuple union) and a single central thing. Prismaticly all metalloenzymes have cheeled metals, usually to ptids or cofactors and protia. And this provides a ring of five members of Pedidos CUC<sub>2</sub>N<sub>2</sub>. Protby or cofactors belong to Metalloenzymes. Organic chelates help to extract metallic ions from rocks or mineral help in hot chemical member. Using chelating metal ions, nutritional supplements are formed. It helps prevent the formation of insoluble complexes in salts for the stomach. In general, these supplements have a greater capacity for the absorption agents common chelars used for the process of softening the dining rooms are Edta and Fósforo. Some frequent Quelada endings include ruthenium chloride with bidental phosphine with bidental phosphine. How sample questions: What are the chelating agents? The effect of the effect of chelato describes the improved affinity of the chelants for a metal ion compared to the affinity of a collection of similar non -scheme (monodentated) ligands for the same metal. Therefore, proteins, polysacians and polynucleic E;cidos are excellent polishing ligands for many metallic ions. Since they do not react with any other element present in the substance. But the natural cheeking agent is mainly used in the cosmã © technical industry. Ethylenediamine teteteethylendiamine as an agent when joining through its two nitrógeno utomos. These molas are generally organic compounds, but this is not a necessity. Describe the origin of the Quelato

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