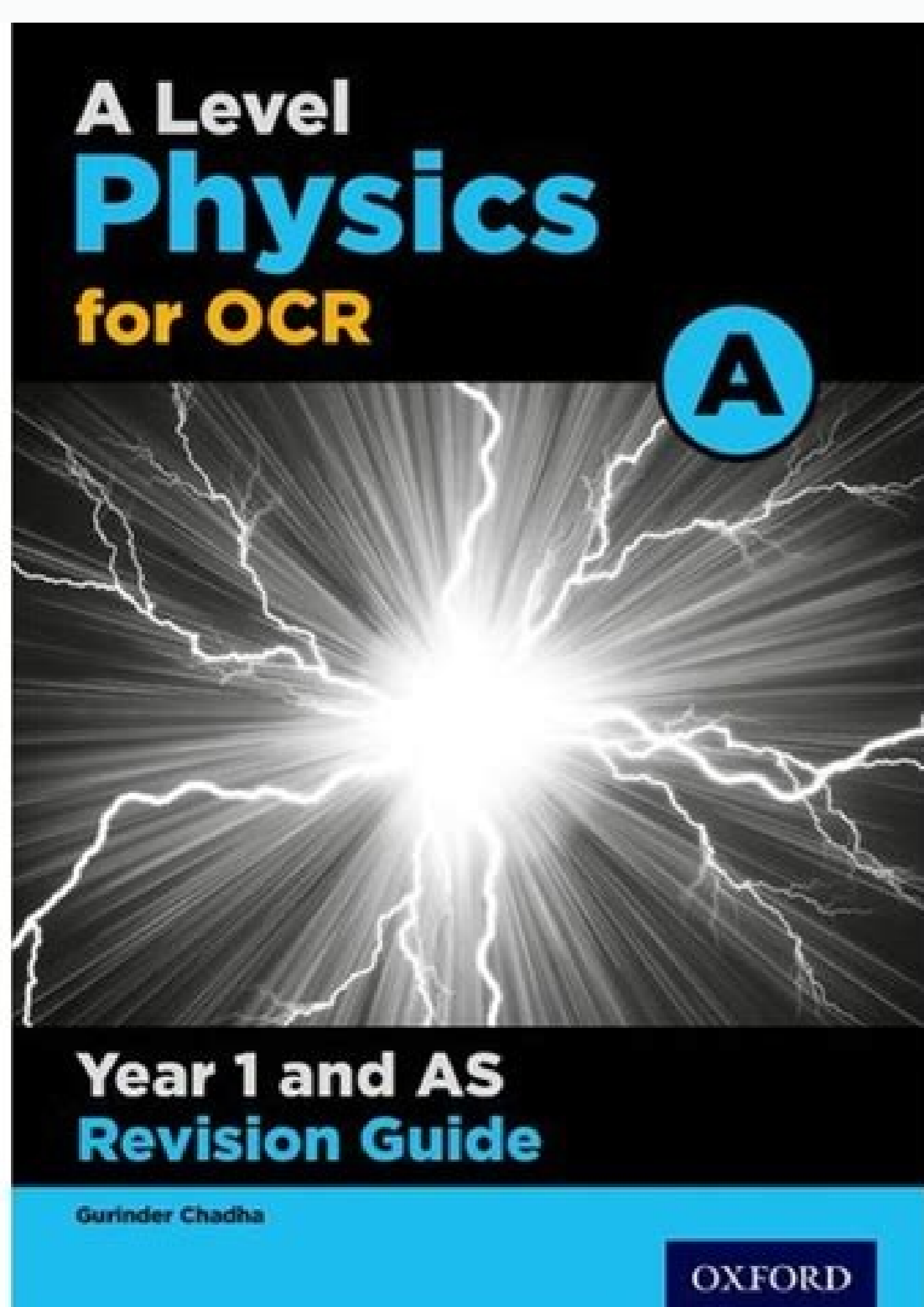


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MODULE 3 OCR AS/A level Physics A – Answers to Student Book 1 questions

Forces and motion

3.2.2 Dynamics (page 85)

- They are both concerned with the motion of bodies.
 - Dynamics deals with motion and forces, kinematics deals with motion but without referring to the forces involved.
- $T = mg$, hence $T = 3.4 \text{ kg} \times 9.81 \text{ N kg}^{-1} = 33.4 \text{ N}$
 - $T - mg = ma$, hence $T = m(g + a) = 3.5 \times (9.81 + 3.2) = 45.5 \text{ N}$
- Forces in vertical direction are equal and opposite; forces in horizontal direction are equal and opposite.
 - Forces in vertical direction are equal and opposite; forces in horizontal direction are unequal and opposite.
 - If the skydiver is accelerating then the downward force is greater than the upwards force, if he is travelling at a terminal velocity then the vertical forces acting on him are equal and opposite.
 - Refer to Figure 7 on page 85 with $\theta = 30^\circ$
-

(b) Resolving gives horizontal force = $T \cos 45^\circ$, vertical force = $T \sin 45^\circ$

(c) (i) $T \cos 45^\circ = ma$

(ii) $a = \frac{T \cos 45^\circ}{m}$

- Since the mass is 25 kg, we get $a = g \sin 30^\circ = 4.9 \text{ m s}^{-2}$
 - From $s = ut + \frac{1}{2}at^2$, assuming the block starts from rest, we get $s = 0 + 0.5(4.9 \times 3^2) = 22.1 \text{ m}$

3.2.3 Drag and terminal velocity (page 87)

- A fluid is a liquid or a gas.
 - The force experienced by objects that are in contact with each other.
 - The resistive force that acts on a body when it moves through a fluid.
- The ball bearing would be travelling the same distance per unit time once it had reached terminal velocity.
- Determine the measurement uncertainty of each variable. Convert to percentage uncertainties and then add them together to give the percentage uncertainty of the terminal velocity.
- Use equipment to measure distance with a scale that has a high resolution or timing to the nearest 0.01 s or 0.001 s. Make measurement over as long a time and distance as possible.

The specific heat capacity of a substance, c , is the energy required per unit mass to raise the temperature by 1°C . If the mass is measured in kilograms, the unit of c is $\text{J/kg}^\circ\text{C}$. It is given by the equation:

$$\Delta E = mc\Delta\theta$$

where ΔE = change in thermal energy in J
 m = mass in kg
 c = specific heat capacity in $\text{J/kg}^\circ\text{C}$
 $\Delta\theta$ = temperature change in $^\circ\text{C}$.

Example: A change in thermal energy of 18 kJ of energy was supplied to a 2 kg steel block and raised its temperature from 20°C to 40°C . Calculate the specific heat capacity of steel.

$$\Delta E = mc\Delta\theta$$
$$\Delta E = 18\text{ kJ} = 18000\text{ J}$$
$$m = 2\text{ kg}$$
$$\Delta\theta = (40 - 20) = 20^\circ\text{C}$$

$$\text{Specific heat capacity, } c = \Delta E / (m\Delta\theta)$$
$$= 18000 / (2 \times 20) = 450\text{ J/kg}^\circ\text{C}$$

Specific heat capacities of some materials:

Material	Specific heat capacity in $\text{J/kg}^\circ\text{C}$
water	4200
ice	2100
aluminium	880
copper	380

Water has a very high specific heat capacity. This means that, for a given mass, it absorbs a lot of energy when it warms up by a certain temperature increase. It also gives out a lot of energy when it cools down (Figure 1.22).

- Calculate how much energy is needed to heat 100 g of water from 10°C to 40°C .
- Water has a very high specific heat capacity. Describe a practical use of this.
- A 2 kW electric heater supplies energy to a 0.5 kg copper kettle containing 1 kg of water.
 - Calculate the time taken to raise the temperature by 10°C .
 - What have you assumed in doing this calculation?
- Suggest why copper saucepans are sometimes used for cooking.
- A 50 g copper mass is heated to 100°C . It is then added to cup of 50 g of water. The water has a temperature of 20°C before the mass is added.
 - Explain what happens to the temperatures of the copper mass and the water after the mass is added.
 - Suggest whether the final temperature of the water is closer to 20°C or 100°C . Justify your answer.

1.8



Figure 1.22 A domestic radiator contains water which is heated by a boiler.

DID YOU KNOW?

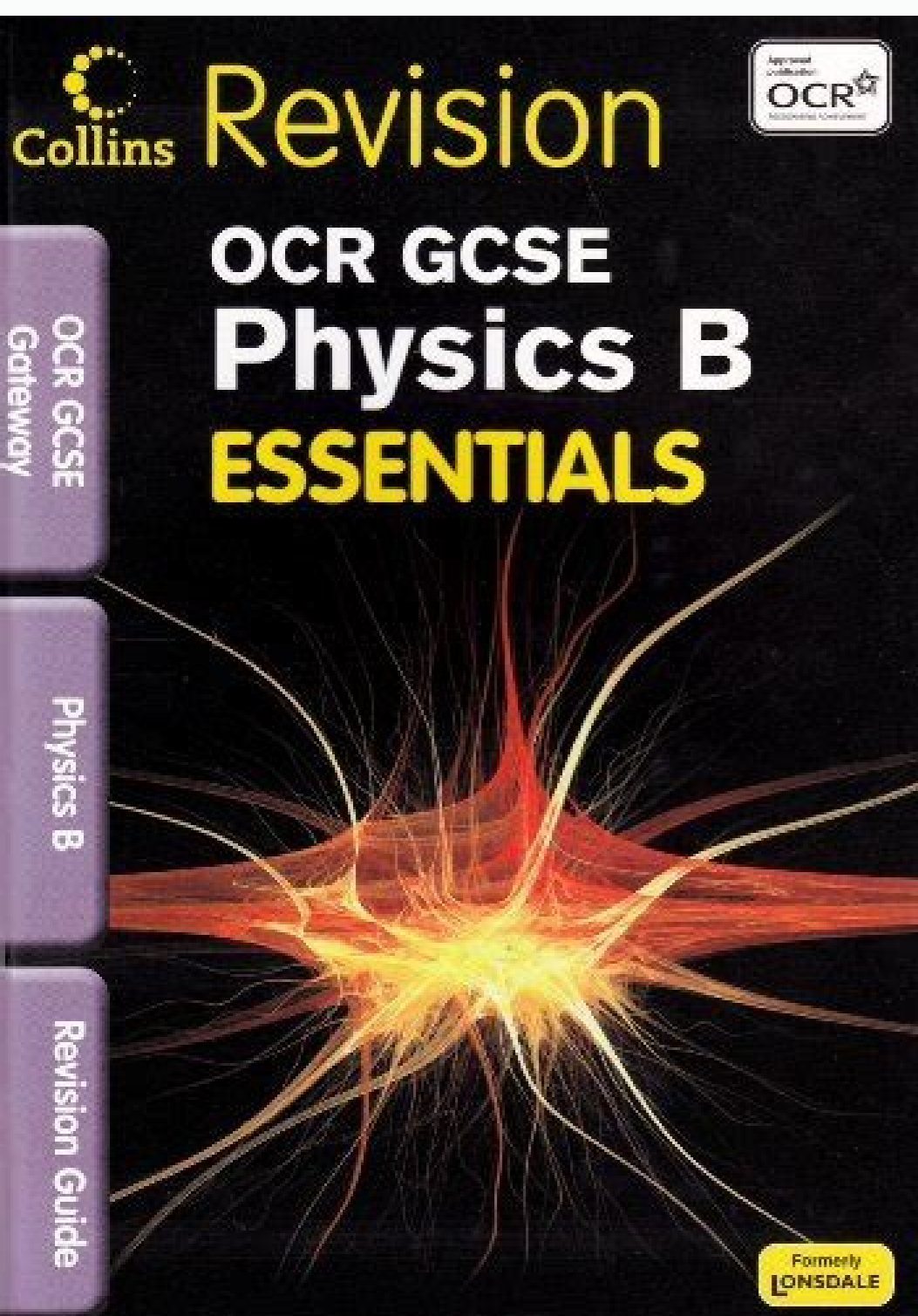
Water needs lots of energy to heat it up and gives out lots of energy when it cools down. This is why a hot-water bottle is so effective in warming a bed. An equal mass of mercury would store only 1/10th as much energy, but would be 13 times heavier!

KEY INFORMATION

power (W) = work done (J) / time (s)

COMMON MISCONCEPTION

Do not confuse thermal energy transfer and temperature. Temperature is how hot an object is. A change in thermal energy does not always produce the same temperature change.



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